



SDI FINAL EVALUATION FORM 1.1

PART 1:

Journal Name:	American Chemical Science Journal
Manuscript Number:	2013 ACSj 4343
Title of the Manuscript:	Thermodynamics of the solvation of lead nitrate in mixed DMF-H2O solvents at 301.15 K.

PART 2:

FINAL EVALUATOR'S comments on revised paper (if any)	Authors' response to final evaluator's comments
<p>I regret to have to conclude that the revised paper by Goma on "Thermodynamics of the solvation of lead nitrate in mixed DMF-H2O solvents at 301.15 K" should not be published in American Chemical Science Journal.</p> <p>1) The English is still bad. 2) The Debye-Hückel expression (1) is inapplicable at the concentrations of > 1 mol/L encountered here. Furthermore, no values for the parameter r° in the denominator were provided. Anyway, the pK_{sp} values in Table 1 in the revised paper are the same as in the previous version, hence were not re-calculated with the new activity coefficient expression! 3) There is no justification provided for the splitting of the Gibbs energy of solvation between the ions according to their ionic radii. 4) Although "the proton solvation free energies in pure water and DMF were -1053 and 1099 KJmol^{-1}, respectively" the entries in the last column of Table 2 are between -1052 and -1056. Hence the authors did not use their own data correctly. 5)etc.</p> <p>It is not necessary to go into detail of all the other misstatements in the paper in order to recommend rejection of this paper.</p>	

Note: Anonymous Reviewer